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PALLADIUM(II) TETRAKIS(1-PYRAZOLYL)BORATE AND SOME  
RELATED SPECIES (U) KENTUCKY UNIV LEXINGTON DEPT OF  
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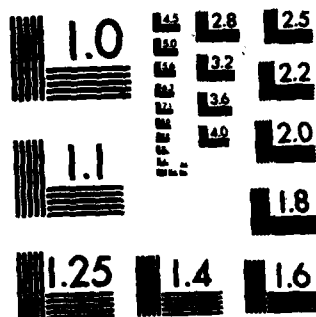
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## REPORT DOCUMENTATION PAGE

unclassified			1b. RESTRICTIVE MARKINGS		
2a. SECURITY CLASSIFICATION AUTHORITY			3. DISTRIBUTION / AVAILABILITY OF REPORT prepared for publication in INORGANICA CHIMICA ACTA		
2b. DECLASSIFICATION / DOWNGRADING SCHEDULE					
4. PERFORMING ORGANIZATION REPORT NUMBER(S) UK/DC/TR-20			5. MONITORING ORGANIZATION REPORT NUMBER(S)		
6a. NAME OF PERFORMING ORGANIZATION Department of Chemistry University of Kentucky		6b. OFFICE SYMBOL (If applicable)	7a. NAME OF MONITORING ORGANIZATION Office of Naval Research		
6c. ADDRESS (City, State, and ZIP Code) Lexington, KY 40506-0055		7b. ADDRESS (City, State, and ZIP Code) 800 North Quincy Street Arlington, VA 22217			
8a. NAME OF FUNDING / SPONSORING ORGANIZATION Office of Naval Research		8b. OFFICE SYMBOL (If applicable)	9. PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER N00014-83-K-0611		
8c. ADDRESS (City, State, and ZIP Code) Arlington, VA 22217		10. SOURCE OF FUNDING NUMBERS			
		PROGRAM ELEMENT NO.	PROJECT NO.	WORK R & T Code 4135003-1	WORK UNIT ACCESSION NO.
11. TITLE (Include Security Classification) PALLADIUM(II) TETRAKIS(1-PYRAZOLYL)BORATE AND SOME RELATED SPECIES					
12. PERSONAL AUTHOR(S) M. K. Das, K. Niedenzu, S. Roy					
13a. TYPE OF REPORT interim technical		13b. TIME COVERED FROM TO		14. DATE OF REPORT (Year, Month, Day) 88/6	
15. PAGE COUNT 8					
16. SUPPLEMENTARY NOTATION					
17. COSATI CODES			18. SUBJECT TERMS (Continue on reverse if necessary and identify by block number)		
FIELD	GROUP	SUB-GROUP	palladium(II) tetrakis(1-pyrazolyl)borate pyrazolylborate/halide complexes		
19. ABSTRACT (Continue on reverse if necessary and identify by block number)					
<p>The structure of the complex <math>\text{Pd}(\text{B}(\text{pz})_4)_2</math> (Hpz = pyrazole) has been explained on the basis of variable-temperature <math>^1\text{H}</math> NMR studies. Reactions of <math>\text{M}[\text{B}(\text{pz})_4]_2</math> (M = Pd, Zn, Pb) with metal dichlorides have been studied and the complexes <math>\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{PdCl}_2</math> and <math>\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{ZnCl}_2</math> have been isolated.</p> <p>(Keywords: Palladium compounds, borates)</p>					
<div style="border: 1px solid black; padding: 5px; display: inline-block;"> <b>DISTRIBUTION STATEMENT A</b>          Approved for public release          Distribution Unlimited       </div> <div style="margin-left: 20px;"> <b>DTIC ELECTE</b>  <b>S JUN 13 1988 D</b>          and       </div>					
20. DISTRIBUTION / AVAILABILITY OF ABSTRACT <input checked="" type="checkbox"/> UNCLASSIFIED/UNLIMITED <input type="checkbox"/> SAME AS RPT. <input type="checkbox"/> DTIC USERS			21. ABSTRACT SECURITY CLASSIFICATION unclassified		
22a. NAME OF RESPONSIBLE INDIVIDUAL			22b. TELEPHONE (Include Area Code)		22c. OFFICE SYMBOL

# Palladium(II) Tetrakis(1-pyrazolyl)borate and Some Related Species

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## Abstract

The structure of the complex  $\text{Pd}[\text{B}(\text{pz})_4]_2$  (Hpz = pyrazole) has been explained on the basis of variable-temperature  $^1\text{H}$  NMR studies. Reactions of  $\text{M}[\text{B}(\text{pz})_4]_2$  ( $\text{M} = \text{Pd}, \text{Zn}, \text{Pb}$ ) with metal dichlorides have been studied and the complexes  $\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{PdCl}_2$  and  $\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{ZnCl}_2$  have been isolated.

## Introduction

Several chain-type polynuclear spiro species containing boron [1] as well as boron and metal [2] centers have been described recently. The general building principle for these materials involved the conversion of a terminal  $\text{B}(\text{pz})_2$  unit (Hpz = pyrazole) into bridging moieties of the type  $\text{B}(\mu\text{-pz})_2\text{M}$  ( $\text{M} = \text{boron or metal}$ ). The complexing  $\text{B}(\text{pz})_2$  unit originated either from poly(1-pyrazolyl)pyrazaboles or from the  $[\text{B}(\text{pz})_4]^-$  ion. The latter was normally employed as the potassium salt, but  $\text{Zn}[\text{B}(\text{pz})_4]_2$  was used in one instance, where the reaction with  $\text{ZnCl}_2$  was found to give a complex  $\text{Zn}[\text{B}(\text{pz})_4]_2 \cdot 2\text{ZnCl}_2$ , which was formulated as the chain-type polynuclear species



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$\text{Cl}_2\text{Zn}(\mu\text{-pz})_2\text{B}(\mu\text{-pz})_2\text{Zn}(\mu\text{-pz})_2\text{B}(\mu\text{-pz})_2\text{ZnCl}_2$ . This result suggested that similar species should exist in which even two different metals are incorporated into the chain. The present report describes data on some metal tetrakis(1-pyrazolyl)borates,  $\text{M}[\text{B}(\text{pz})_4]_2$ , and their interaction with metal dichlorides,  $\text{M}'\text{Cl}_2$ .

### Results and Discussion

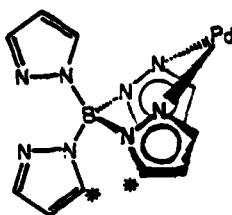
**Starting Materials.** The three species  $\text{M}[\text{B}(\text{pz})_4]_2$  with  $\text{M} = \text{Zn}, \text{Pb},$  and  $\text{Pd}$  were selected as starting materials for the present study. The  $^1\text{H}$  NMR spectra of the Zn [2] and the Pb species are quite similar exhibiting only three sharp signals in 1:1:1 ratio. Essentially no changes were observed in the spectra over the temperature range from +25 to -50 °C indicating the equivalence of all four pz groups. This is essentially the same fluxional situation as that observed for other metal tetrakis(1-pyrazolyl)borates but is significantly different from the  $^1\text{H}$  NMR spectrum of  $\text{Pd}[\text{B}(\text{pz})_4]_2$ . This latter compound has been mentioned in the literature [3] but no experimental data were given. It is, however, easily prepared from  $\text{M}_2[\text{PdCl}_4]$  ( $\text{M} = \text{Na}, \text{K}$ ) and  $\text{M}'[\text{B}(\text{pz})_4]$  ( $\text{M}' = \text{Li}, \text{K}$ ), although the purification is somewhat tedious.

In contrast to the zinc and lead salts cited above, the palladium complex exhibits a much more complicated  $^1\text{H}$  NMR spectrum. At -48 °C the spectrum shows sets of signals for three different pz groups with  $\delta(^1\text{H})$  7.86/6.93/6.36, 7.79/7.38/6.46, and 7.41/7.09/6.32 in a 1:1:2 intensity ratio, of which the two former are merging into a single set near room temperature (see Figure 1). A variable-temperature study over the range from

Fig. 1

-25 to +80 °C (performed in  $\text{CDCl}_2\text{CDCl}_2$ ) showed that the lowest field signals of the two sets of intensity 1 merge near 40 °C, the highest field signals merge near 60 °C, but the central set merges only near 70 °C. The data suggest that the species should be formulated as  $(\text{pz})_2\text{B}(\mu\text{-pz})_2\text{Pd}(\mu\text{-pz})_2\text{B}(\text{pz})_2$ . This is based on the following interpretation. In the low-temperature spectrum, the  $^1\text{H}$  NMR signals of the presumably puckered  $\text{B}(\mu\text{-pz})_2\text{Pd}$  rings are assigned to the set  $\delta(^1\text{H})$  7.41/7.09/6.32. This signal set is sharp over the entire temperature range studies and these pz groups remain in the same magnetic environment regardless of the temperature, but they obviously affect the terminal boron-bonded pz groups differently. At low temperatures, the pseudo-axial and -equatorial

terminal boron-bonded pz groups give distinctly separate signals until, with raising of the temperature, inversion of the  $B(\mu\text{-pz})_2\text{Pd}$  ring becomes rapid on the NMR time scale, and, ultimately, the signals of the terminal pz groups collapse to a single set. Assuming that the 5-H of the bottom terminal pz group in 1 (one half of the molecule) will



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be in a magnetic environment most similar to those of the 5-H of the bridging pz groups (both marked with an asterisk in 1), the signal of the former is assigned at 7.38 ppm, those of the latter at 7.41 ppm. This automatically leads to the assignments of all other proton resonance signals in the low-temperature spectrum.

The foregoing data illustrate that in the Pb and Zn complex all pz groups are equivalent in their bonding to the metal. In contrast, the same two pz groups of each  $B(\text{pz})_4$  moiety of the Pd complex are always bonded to the metal. The non-planar  $B(\mu\text{-pz})_2\text{Pd}$  rings are static at low temperatures but begin to invert with increasing temperature; only at high temperatures is the inversion sufficiently fast to provide for equivalence of the terminal pz groups, which do not exchange sites with those bridging between the B and the Pd.

**The Interaction of  $M[B(\text{pz})_4]_2$  with  $M'\text{Cl}_2$ .** Recently, the preparation of the unusual complex  $(\text{pz})_2B(\mu\text{-pz})_2\text{Pd}(\mu\text{-pz})_2B(\mu\text{-pz})_2\text{PdCl}_2 \cdot 0.5\text{CH}_2\text{Cl}_2$  has been reported [4]. This compound would be an interesting starting material for chain elongation at the remaining terminal  $B(\text{pz})_2$  site. Unfortunately, all attempts to duplicate the reported preparation of this material failed and only  $\text{Pd}[B(\text{pz})_4]_2$  was isolated when following the original directions. On the other hand, reaction of  $\text{Pd}[B(\text{pz})_4]_2$  with  $\text{PdCl}_2$  gave the complex  $\text{Pd}[B(\text{pz})_4]_2 \cdot 2\text{PdCl}_2$ , independent of the stoichiometry of the reagents. In no case was a species of the composition  $\text{Pd}[B(\text{pz})_4]_2 \cdot \text{PdCl}_2$  obtained.

The complex  $\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{PdCl}_2$  is insoluble in most common solvents but NMR spectra were obtained on a solution in  $\text{DMSO}-d_6$ . The  $^1\text{H}$  NMR spectrum of the material is extremely complex and suggests the presence of several types (at least five!) of pz groups, based on the number of pz 4-H signals. The  $^{11}\text{B}$  NMR spectrum exhibits a sharp signal at 0.4 ppm but with a clearly recognizable and intense shoulder at -0.1 ppm. These observations suggest exchange reactions to occur in solution but which could not be specified: A variable-temperature  $^1\text{H}$  NMR study was inconclusive.

The reaction of  $\text{Zn}[\text{B}(\text{pz})_4]_2$  with  $\text{ZnCl}_2$  has previously been reported to give a species which was formulated as  $\text{Cl}_2\text{Zn}(\mu\text{-pz})_2\text{B}(\mu\text{-pz})_2\text{Zn}(\mu\text{-pz})_2\text{B}(\mu\text{-pz})_2\text{ZnCl}_2$  [2].  $\text{Pb}[\text{B}(\text{pz})_4]_2$  does not react with  $\text{PbCl}_2$  under similar conditions. However, the reaction of  $\text{Pd}[\text{B}(\text{pz})_4]_2$  with  $\text{ZnCl}_2$  was found to give the complex  $\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{ZnCl}_2$ , which probably can also be formulated as  $\text{Cl}_2\text{Zn}(\mu\text{-pz})_2\text{B}(\mu\text{-pz})_2\text{Pd}(\mu\text{-pz})_2\text{B}(\mu\text{-pz})_2\text{ZnCl}_2$ .

Surprisingly, when  $\text{Zn}[\text{B}(\text{pz})_4]_2$  was reacted with  $\text{PdCl}_2$ , the complex  $\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{PdCl}_2$  was obtained rather than the desired mixed-metal species. Reaction of  $\text{Zn}[\text{B}(\text{pz})_4]_2$  with  $\text{MCl}_2$  ( $\text{M} = \text{Co}, \text{Ni}$ ; stirring the reagents in 1:2 molar ratio in ether at room temperature and recrystallizing the insoluble material from toluene) was found to give species of the elemental composition  $2\text{Zn}[\text{B}(\text{pz})_4]_2 \cdot 3\text{M}[\text{B}(\text{pz})_4]_2$  (the  $\text{M} = \text{Ni}$  complex is pink, and the one with  $\text{M} = \text{Co}$  is yellow). The  $^{11}\text{B}$  NMR spectra exhibited a broad (ca. 4000 Hz) base peak (centered near -2 ppm) with a sharp spike at -0.4 ppm, but these species were not yet further investigated. Reaction of  $\text{Pd}[\text{B}(\text{pz})_4]_2$  with  $\text{NiCl}_2$  gave  $\text{Ni}[\text{B}(\text{pz})_4]_2$ , but no reaction was observed with  $\text{CoCl}_2$  (mixtures in DMF, 40 h stirring at room temperature).

#### Experimental Section

Elemental analyses were performed by the Schwarzkopf Microanalytical Laboratory, Woodside, NY. Melting points (uncorrected) were determined on a Mel-Temp block.

NMR spectra were recorded on a Varian XL-200 instrument. Chemical shift data are given in ppm with positive values indicating downfield from the reference (internal  $\text{Me}_4\text{Si}$  for  $^1\text{H}$  NMR and external  $\text{Et}_2\text{O} \cdot \text{BF}_3$  for  $^{11}\text{B}$  NMR): s = singlet, d = doublet, t = triplet, q = quartet, p = quintuplet, m = unresolved multiplet, and an asterisk

denotes a broad signal. Coupling constants  $J$  are given in Hz. Details for 2D NMR experiments have been given elsewhere [5].

$\text{Pd}[\text{B}(\text{pz})_4]_2$ . Solid  $\text{K}[\text{B}(\text{pz})_4]$  [3] (18.44 g, 58 mmol) was added in small portions to a stirred aqueous solution of  $\text{Na}_2[\text{PdCl}_4]$  (10 g of the trihydrate, 28.7 mmol). An immediate precipitate formed and the mixture was stirred at room temperature for 18 h. The precipitate was collected, washed with water and dried in air. The material was purified by repeated dissolution in  $\text{CH}_2\text{Cl}_2$  and careful (slow) precipitation with diethyl ether to yield 8.5 g (45%) of purified pale yellow material, which begins to decompose near 350 °C. Anal. Calcd for  $\text{C}_{24}\text{H}_{24}\text{B}_2\text{N}_{16}\text{Pd}$  ( $M_r$  664.59): C, 43.37; H, 3.64; N, 33.71. Found: C, 43.41; H, 3.81; N, 32.16.

NMR data (solution in  $\text{CD}_2\text{Cl}_2$ ):  $\delta$  ( $^1\text{H}$ ) 7.77\* (2 H), 7.40 (2 H, d,  $J = 2.7$ , of d,  $J = 0.7$ ), 7.33\* (1 H), 7.05 (2 H, d,  $J = 2.1$ , of d,  $J = 0.7$ ), 6.96\* (1 H), 6.42\* (1 H), 6.33\* (1 H), 6.27 (2 H, t,  $J = 2.4$ );  $\delta$  ( $^{11}\text{B}$ ) 0.9 ( $h_{1/2} = 40$  Hz). At -48 °C:  $\delta$  ( $^1\text{H}$ ) 7.86 (1 H, d,  $J = 1.6$ ), 7.79 (1 H, d,  $J = 1.5$ ), 7.41 (2 H, d,  $J = 2.6$ ), 7.38 (1 H, d,  $J = 2.3$ ), 7.09 (2 H, d,  $J = 2.2$ , of d,  $J = 0.7$ ), 6.93 (1 H, d,  $J = 2.4$ ), 6.46 (1 H, two overlapping d), 6.36 (1 H, two overlapping d), 6.32 (2 H, t,  $J = 2.4$ ). As is based on low-temperature 2D-NMR data, the sets  $\delta$  ( $^1\text{H}$ ) 7.86/6.93/6.36, 7.79/7.38/6.46, and 7.41/7.09/6.32, respectively, belong to individual pz groups. Solution in  $\text{CDCl}_3$  at 25 °C:  $\delta$  ( $^1\text{H}$ ) 7.84\* (1 H, unresolved), 7.79\* (1 H, unresolved), 7.44 (2 H, d,  $J = 2.5$ ), 7.30\* (1 H, unresolved), 7.06\* (1 H, unresolved), 7.05 (2 H, d,  $J = 2.2$ , of d,  $J = 0.7$ ), 6.49\* (1 H, unresolved), 6.33\* (1 H, unresolved), 6.26 (2 H, t,  $J = 2.3$ );  $\delta$  ( $^{11}\text{B}$ ) 0.9 ( $h_{1/2} = 40$  Hz). At 40 °C:  $\delta$  ( $^1\text{H}$ ) 7.80\* (2 H), 7.43 (2 H), 7.28\* (1 H), 7.05 (3 H, unresolved with a broad baseline), 6.36\* (2 H), 6.25 (2 H, unresolved).

$\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{ZnCl}_2$ . A solution of 1.25 g (1.9 mmol) of  $\text{Pd}[\text{B}(\text{pz})_4]_2$  in 50 mL of  $\text{CH}_2\text{Cl}_2$  was added dropwise with stirring to a solution of 3.8 mmol of  $\text{ZnCl}_2$  in diethyl ether. The mixture was stirred at ambient temperature for 48 h. The insoluble material was collected, washed with ether and air-dried to yield 1.20 g (68%) of pale yellow crystals, mp 326–330 °C decomposition. Anal. Calcd for  $\text{C}_{24}\text{H}_{24}\text{B}_2\text{Cl}_4\text{N}_{16}\text{PdZn}_2$  ( $M_r$  937.2): C, 30.8; H, 2.6; B, 2.3; Cl, 15.1; N, 23.9; Pd, 11.3; Zn, 13.9. Found: C, 30.50; H, 2.72; B, 2.07; Cl, 14.75; N, 23.28; Pd, 10.92; Zn, 13.5.



NMR data (solution in DMF- $d_6$ ):  $\delta$  ( $^1\text{H}$ ) 7.84\* (1 H), 7.76\* (2 H), 7.43 (2 H, d,  $J = 2.5$ , of d,  $J = 0.6$ ), 7.26 (2 H, d,  $J = 2.1$ , of d,  $J = 0.6$ ), 6.92\* (1 H), 6.46 (2 H, t,  $J = 2.4$ ), 6.40\* (2 H);  $\delta$  ( $^{11}\text{B}$ ): 1.3 ( $h_{1/2} = 65$  Hz). At  $-25^\circ\text{C}$ :  $\delta$  ( $^1\text{H}$ ) 7.94 (1 H, d,  $J = 1.1$ ), 7.92 (1 H, d,  $J = 2.3$ ), 7.83 (1 H, d,  $J = 1.3$ ), 7.45 (2 H, d,  $J = 2.5$ ), 7.33 (2 H, d,  $J = 1.6$ ), 6.90 (1 H, d,  $J = 2.0$ ), 6.53 (2 H, t,  $J = 2.4$ ), 6.50 (1 H, unsym t = two overlapping d,  $J$  ca. 1.6), 6.44 (1 H, unsym t = two overlapping d,  $J$  ca. 1.6). As based on selective decoupling experiments, the sets  $\delta$  ( $^1\text{H}$ ) 7.94/6.90/6.44, 7.92/7.83/6.50, and 7.45/7.33/6.53 (1:1:2 area ratio) belong to individual pz groups. At  $75^\circ\text{C}$ :  $\delta$  ( $^1\text{H}$ ) 7.78 (2 H, unresolved), 7.44 (2 H, unresolved), 7.3\* (2 H), 7.22 (2 H, unresolved), 6.41 (4 H, unresolved).

$\text{Pd}[\text{B}(\text{pz})_4]_2 \cdot 2\text{PdCl}_2$ . Solid  $\text{PdCl}_2$  (0.54 g, 3.0 mmol) was added in small portions to a stirred solution of 1.0 g (1.5 mmol) of  $\text{Pd}[\text{B}(\text{pz})_4]_2$  in 25 mL of  $\text{CH}_2\text{Cl}_2$ . The mixture was stirred at room temperature for 48 h. The yellow-brown precipitate was collected, washed with  $\text{CH}_2\text{Cl}_2$  and dried under vacuum. After recrystallization from dimethylformamide an apricot-colored material (64.5% yield based on  $\text{Pd}[\text{B}(\text{pz})_4]_2$ ) was obtained, which begins to decompose near  $320^\circ\text{C}$ . Anal. Calcd for  $\text{C}_{24}\text{H}_{24}\text{B}_2\text{Cl}_4\text{N}_{16}\text{Pd}_3$  ( $M_r$  1019.2): C, 28.29; H, 2.37; B, 2.12; Cl, 13.91; N, 21.99; Pd, 31.32. Found: C, 28.41; H, 2.45; B, 2.10; Cl, 13.68; N, 22.20; Pd, 31.13.

NMR data (solution in DMSO- $d_6$ ):  $\delta$  ( $^1\text{H}$ ) see text;  $\delta$  ( $^{11}\text{B}$ ) 0.4 (s,  $h_{1/2} = 50$  Hz with a sharp shoulder at  $-0.1$ ).

$\text{Pb}[\text{B}(\text{pz})_4]_2$  [3]. NMR data (solution in  $\text{CDCl}_3$ ):  $\delta$  ( $^1\text{H}$ ) 7.62 (1 H, d,  $J = 1.9$ , of d,  $J = 0.6$ ), 7.39 (1 H, d,  $J = 2.4$ , of d,  $J = 0.5$ ), 6.33 (1 H, unsymmetrical t = 2 overlapping d);  $\delta$  ( $^{11}\text{B}$ ) 2.0 (s,  $h_{1/2} = 30$  Hz). At  $-50^\circ\text{C}$  a slight broadening of the two low-field  $^1\text{H}$  NMR signals is observed but the high-field signal remains undisturbed.

**Acknowledgment.** This work was supported by the U. S. Office of Naval Research (K.N.) and the Department of Science and Technology, India (M.K.D.).

## References

1. C. M. Clarke, K. Niedenzu, P. M. Niedenzu and S. Trofimenko, *Inorg. Chem.*, **24**, 2648 (1985).
2. J. Bielawski, T. G. Hodgkins, W. J. Layton, K. Niedenzu, P. M. Niedenzu and S. Trofimenko, *Inorg. Chem.*, **25**, 87 (1986).
3. S. Trofimenko, *J. Am. Chem. Soc.*, **89**, 3170 (1967).

4. M. Onishi, K. Hiraki, A. Ueno, Y. Yamaguchi and Y. Ohama, *Inorg. Chim. Acta*, 82, 121 (1984).
5. W. J. Layton, K. Niedenzu and S. L. Smith, *Z. Anorg. Allg. Chem.*, 495, 52 (1982).

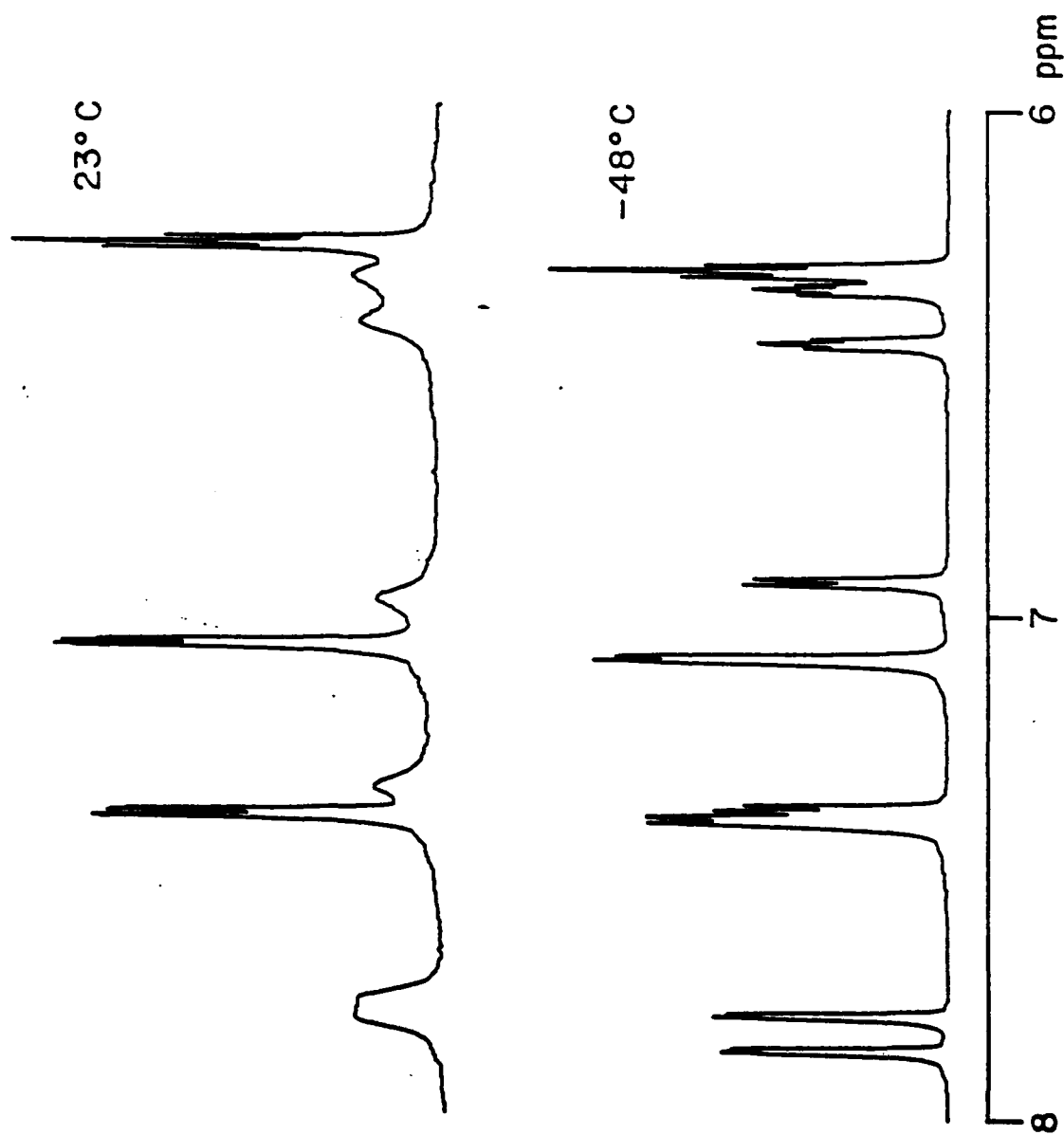


Fig. 1. The  $^1\text{H}$  NMR spectrum of  $\text{Pd}[\text{B}(\text{pz})_4]_2$  in  $\text{CD}_2\text{Cl}_2$  at  $+23$  and  $-48^\circ\text{C}$ .

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